The Periodic Table Lesson 12 - Group 7 Displacement

Science

Chemistry - Key Stage 3

Miss Willett



What have you learnt already?

Where are the alkali metals on the periodic table? ٦.

Where are the halogens on the periodic table? 2.

What is the charge of an electron? 3.



What is a displacement reaction?

Write a definition for a displacement reaction:

Think: REACTIVE, PLACE, COMPOUND



True or false?

Chlorine is the most reactive halogen

The nucleus of an atom is negative

Electrons are easier to gain for small atoms

Smaller halogens are more reactive



Group 7 reactivity

Explain why reactivity decreases down Group 7:

Atoms at the top of the group are..

The nucleus is...

The atom is trying to..

So the close the outer shell is to..

Think: Charge of nucleus? Lose/gain electrons? Smaller / bigger atoms?

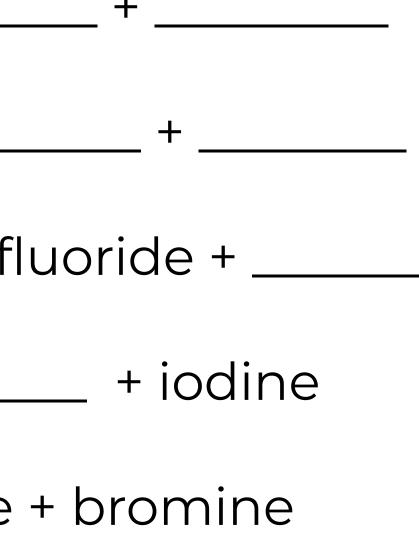
Correct me!

Find my three mistakes this equation:

lodine + sodium fluoride = sodium iodide + fluoride



Reactions of group 7 + group 1 Complete the following word equations: Bromine + calcium iodide → _____ + ____ Chlorine + sodium bromide → _____ + ____ Lithium chloride + _____ → _____ fluoride + _____ _____ + bromine → caesium _____ + iodine • _____ + _____ \rightarrow sodium fluoride + bromine





Finishing off. Predict which reactions will take place (\checkmark or x) Write the products in the box where Chlorine Bromin Potassium chloride Potassium bromide Potassium iodide

е	lodine