

Structures and bonding

Further Ionic Bonding

Worksheet

Combined science - Chemistry - Key Stage 4

Mr Robbins



Periodic Table of Elements

Key:

relative atomic mass

1

H

hydrogen

1

Atomic symbol

Atomic (proton number)

1 H hydrogen 1	4 He helium 2																
7 Li lithium 3	9 Be beryllium 4																
23 Na sodium 11	24 Mg magnesium 12																
11 B boron 5	12 C carbon 6	14 N nitrogen 7	16 O oxygen 8	19 F fluorine 9	20 Ne neon 10												
27 Al aluminium 13	28 Si silicon 14	31 P phosphorus 15	32 S sulfur 16	35.5 Cl chlorine 17	40 Ar argon 18												
39 K potassium 19	40 Ca calcium 20	45 Sc scandium 21	48 Ti titanium 22	51 V vanadium 23	52 Cr chromium 24	55 Mn manganese 25	56 Fe iron 26	59 Co cobalt 27	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	70 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	80 Br bromine 35	84 Kr krypton 36
85 Rb rubidium 37	88 Sr strontium 38	89 Y yttrium 39	91 Zr zirconium 40	93 Nb niobium 41	96 Mo molybdenum 42	[97] Tc technetium 43	101 Ru ruthenium 44	103 Rh rhodium 45	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	128 Te tellurium 52	127 I iodine 53	131 Xe xenon 54
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	178 Hf hafnium 72	181 Ta tantalum 73	184 W tungsten 74	186 Re rhenium 75	190 Os osmium 76	192 Ir iridium 77	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[267] Rf rutherfordium 104	[270] Db dubnium 105	[269] Sg seaborgium 106	[270] Bh bohrium 107	[270] Hs hassium 108	[278] Mt meitnerium 109	[281] Ds darmstadtium 110	[281] Rg roentgenium 87	[285] Cn copernicium 112	[286] Nh nihonium 113	[289] Fl flerovium 114	[289] Mc moscovium 115	[293] Lv livermorium 116	[293] Ts tennessine 117	[294] Og oganeson 118

* The lanthanides (atomic numbers 58 - 71) and the Actinides (atomic numbers 90 - 103) have been omitted.

Relative atomic masses for **Cu** and **Cl** have not been rounded to the nearest whole number.



Draw a diagram to show what happens when Calcium (Ca) Reacts with Chlorine (Cl)

- 1. Show the atoms outside shells**
- 2. Draw arrows to show how the electrons move**
- 3. Draw the ions with brackets and charges**
- 4. Ca is in group 2 and Cl is in group 7**



Giant ionic lattice

Ionic compounds form giant ionic lattices

- Giant = Big
- Ionic = Made of oppositely charged ions
- Lattice = Repeating structure

E.g. NaCl



Independent practice:

Copy and complete the table. The first one has been completed for you.

Name	Positive ion	Negative ion	Formula
sodium oxide	Na ⁺	O ²⁻	Na ₂ O
Magnesium oxide			
Potassium iodide			
Magnesium chloride			
Aluminium chloride			
Calcium chloride			
Calcium oxide			

Hint	
Group No.	Ion charge
1	+1
2	+2
3	+3
6	-2
7	-1



Your turn

Magnesium Oxide contains Magnesium (Mg^{2+}) ions and Oxide (O^{2-}) ions.

Describe in terms of electrons what happens when Magnesium reacts

- Change in the metal
- Change in the non-metal
- Attraction between opposite charges
- Ionic formula



Your Turn

Magnesium Bromide contains Magnesium (Mg^{2+}) ions and Bromide (Br^-) ions.

Describe in terms of electrons what happens when Magnesium reacts with Bromine

- Change in the metal
- Change in the non-metal
- Attraction between opposite charges
- Ionic formula

