

# Reactivity

## Lesson 1 - Electron Configuration

Chemistry - Key Stage 3

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**Magnesium has an atomic number of 12 and a mass number of 24.**

**How many protons does it have?**

**How many electrons does it have?**

**How many neutrons does it have?**





**Magnesium has an atomic number of 12 and a mass number of 24.**

**12 protons**

**12 electrons**

**$24 - 12 = 12$  neutrons**



# Independent Practice

How many protons, neutrons and electrons does each element contain?

<p><b>32</b></p> <p><b>S</b></p> <p>sulfur</p> <p><b>16</b></p>	<p><b>19</b></p> <p><b>F</b></p> <p>fluorine</p> <p><b>9</b></p>	<p><b>23</b></p> <p><b>Na</b></p> <p>sodium</p> <p><b>11</b></p>
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# Independent Practice

**Sulfur:** 16 protons, 16 electrons, 16 neutrons

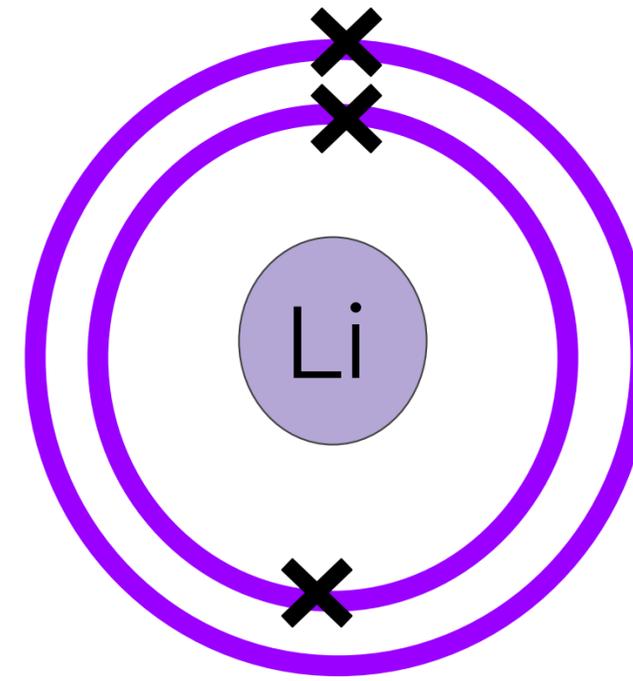
**Fluorine:** 9 protons, 9 electrons, 10 neutrons

**Sodium:** 11 protons, 11 electrons, 12 neutrons



## Independent practice:

1. Draw the electron configuration for lithium which has 3 electrons.
2. Write the electron configuration for boron which has 5 electrons.
3. Write the electron configuration for sodium which has 11 electrons.



2.3

2.8.1





**What does the group number of an element tell us? Can you see a pattern?**

The group number tells us how many electrons are in the outer shell of the atom.

