

Reactivity

Lesson 1 - Electron Configuration

Chemistry - Key Stage 3

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Magnesium has an atomic number of 12 and a mass number of 24.

How many protons does it have?

How many electrons does it have?

How many neutrons does it have?





Magnesium has an atomic number of 12 and a mass number of 24.

12 protons

12 electrons

$24 - 12 = 12$ neutrons



Independent Practice

How many protons, neutrons and electrons does each element contain?

<div>32</div> <div>S</div> <div>sulfur</div> <div>16</div>	<div>19</div> <div>F</div> <div>fluorine</div> <div>9</div>	<div>23</div> <div>Na</div> <div>sodium</div> <div>11</div>
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Independent Practice

Sulfur: 16 protons, 16 electrons, 16 neutrons

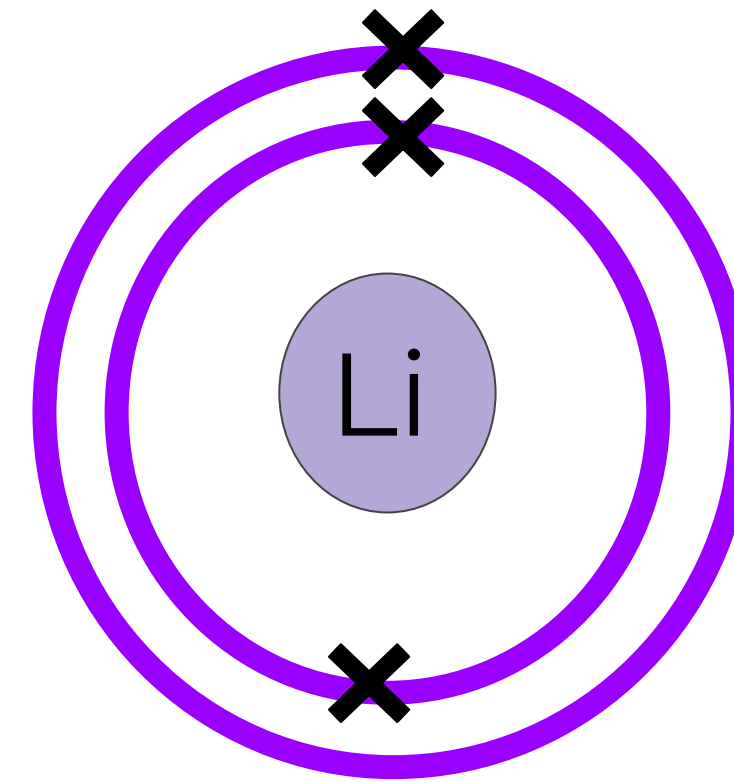
Fluorine: 9 protons, 9 electrons, 10 neutrons

Sodium: 11 protons, 11 electrons, 12 neutrons



Independent practice:


1. Draw the electron configuration for lithium which has 3 electrons.
2. Write the electron configuration for boron which has 5 electrons.
3. Write the electron configuration for sodium which has 11 electrons.



2.3

2.8.1





**What does the group number
of an element tell us? Can
you see a pattern?**

The group number tells us how
many electrons are in the
outer shell of the atom.

