

# Reactivity

## Lesson 7 - Making a Salt

Chemistry - Key Stage 3

Miss Fenner



**Salt is made from oppositely charged ions.  
These ions \_\_\_\_\_ each other.**

**Option 1**

Attract

**Option 2**

Repel



# Which of the following is NOT a common salt type?

Option 1

Sulfate

Option 2

Chloride

Option 3

Nitrate

Option 4

Oxide



# Independent Practice

True or false

1. Salts are always edible and tasty on chips.
2. The proper name for table salt is sodium chloride.
3. Salts are made of ions with the same charge.
4. Salts are made by reacting oxygen with a metal.
5. Common endings on the end of salt names are sulfate, chloride and nitrate.



# Independent Practice

True or false

1. Salts are always edible and tasty on chips.
2. The proper name for table salt is sodium chloride.
3. Salts are made of ions with the same charge.
4. Salts are made by reacting oxygen with a metal.
5. Common endings on the end of salt names are sulfate, chloride and nitrate.



Filtration is a method used to separate a **solid** from a liquid.



Filtration uses equipment including a funnel, filter paper and a beaker.



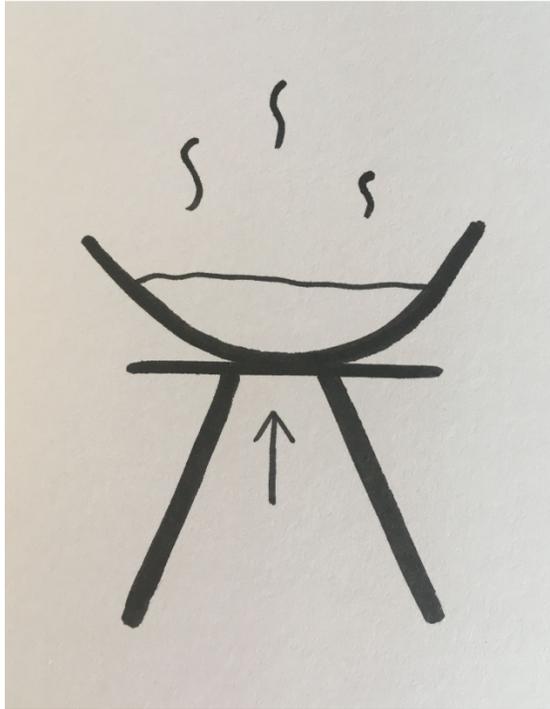
Crystallisation is a method used to separate a **dissolved solid** from a solution.



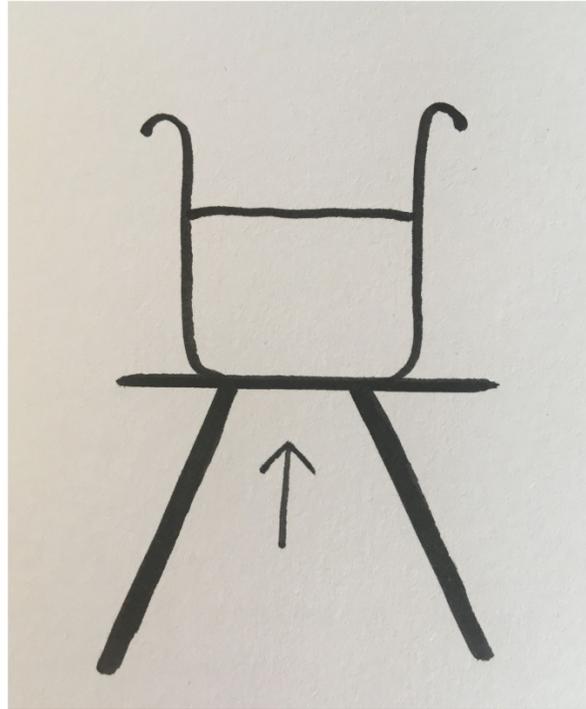
Crystallisation uses equipment including  
a tripod, a bunsen burner and an  
evaporating basin.



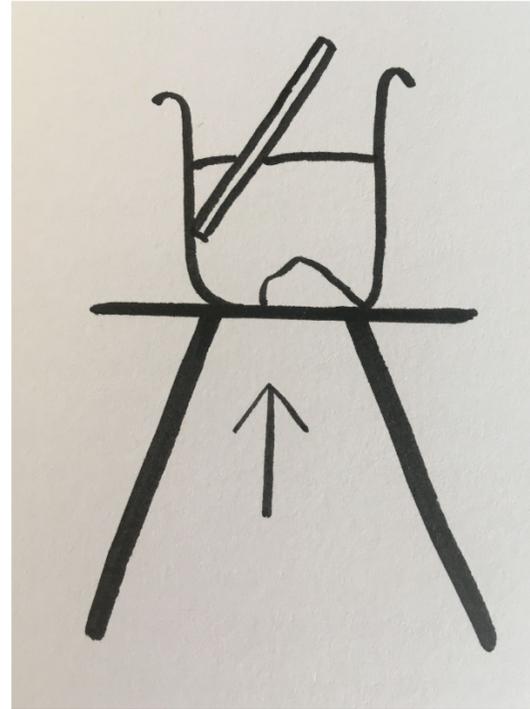
These pictures show each step of the method used to make a salt.  
Put the pictures in the correct order.



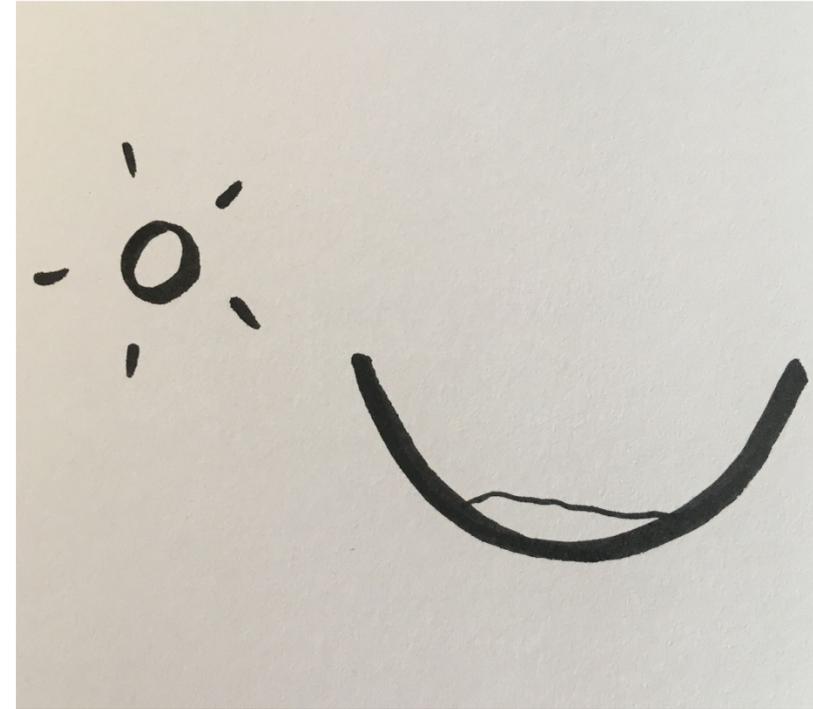
A



B



C



D

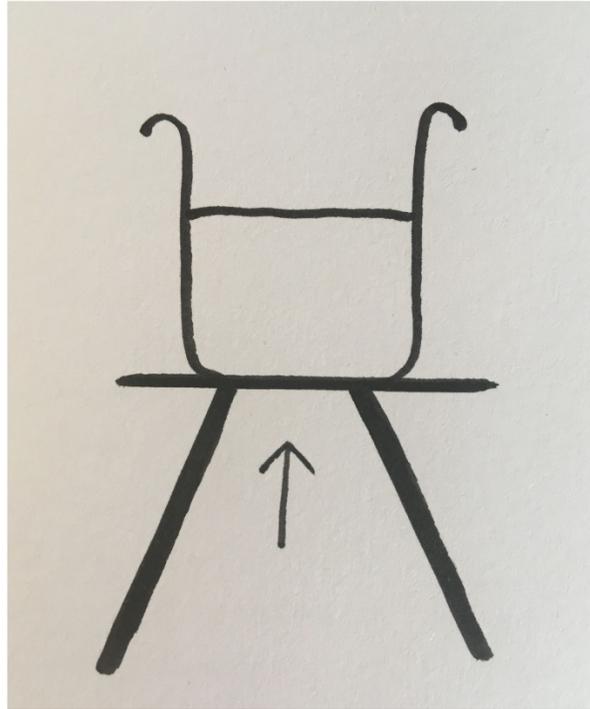


E

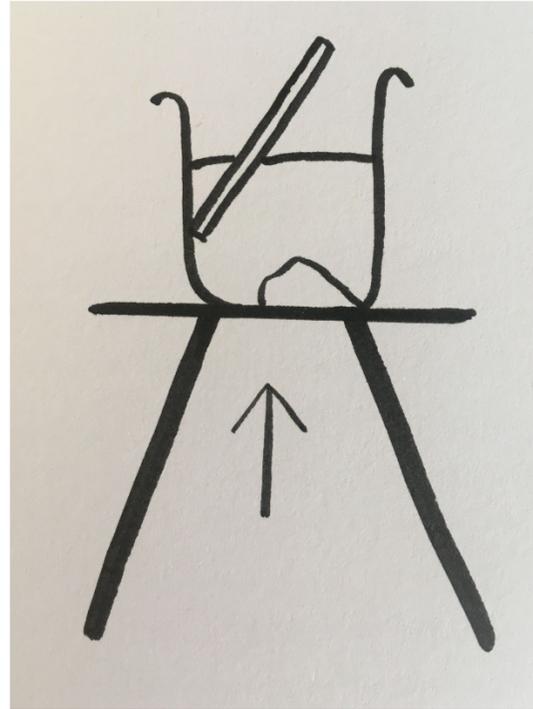
Source: Miss Fenner



These pictures show each step of the method used to make a salt.  
Put the pictures in the correct order.



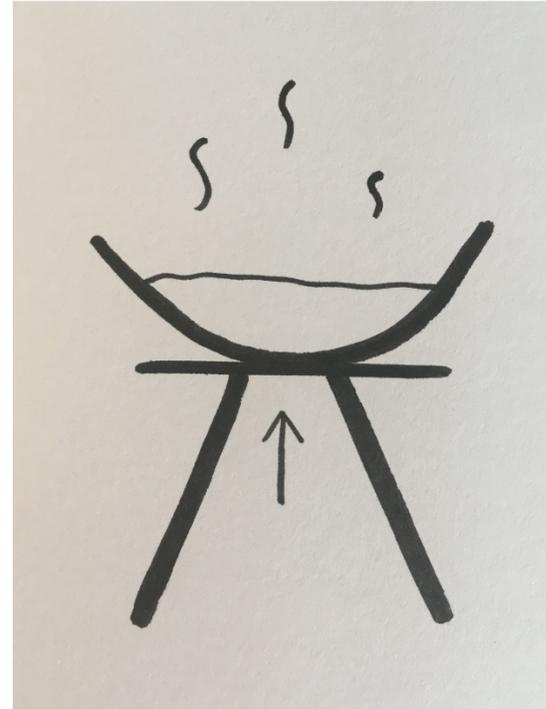
B



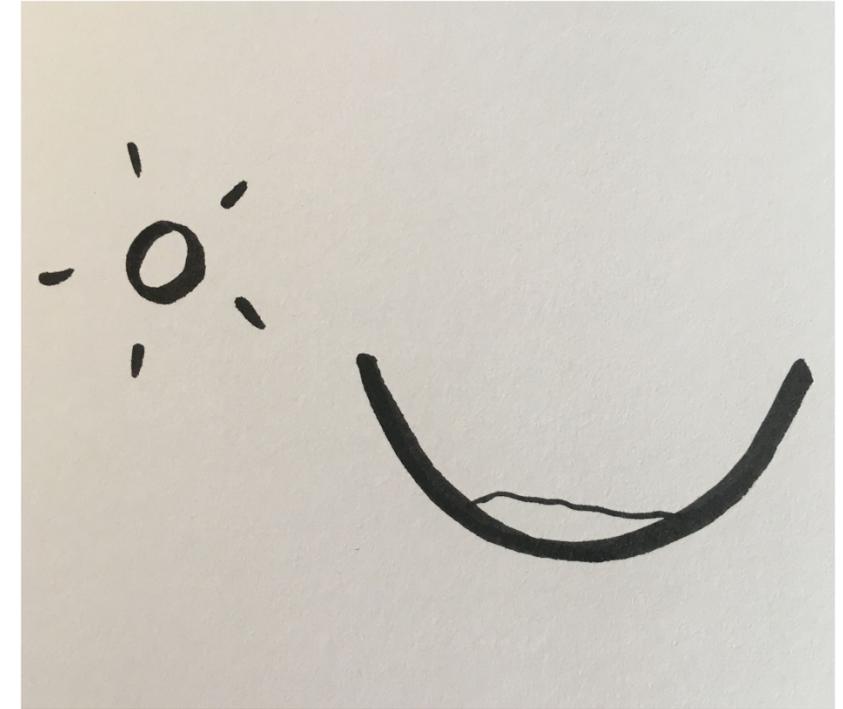
C



E



A

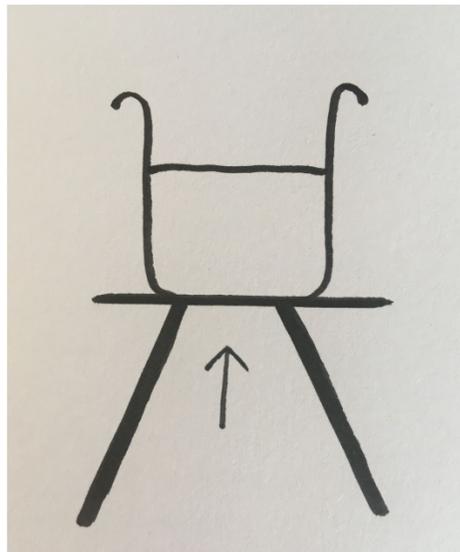


D

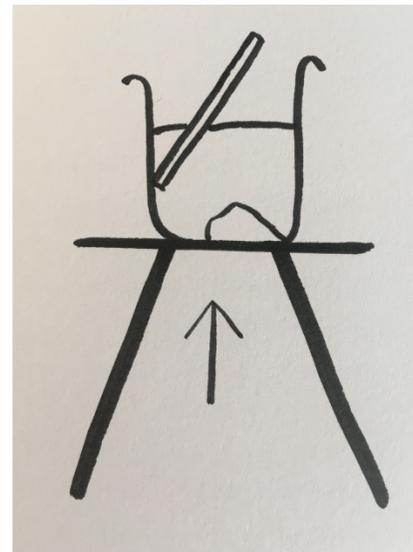


# Independent Practice

Write 1/2 sentence(s) to explain what is happening in each step. Use the diagram to help you.



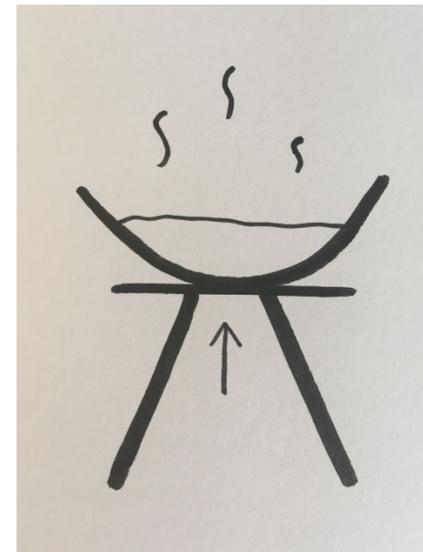
Step 1



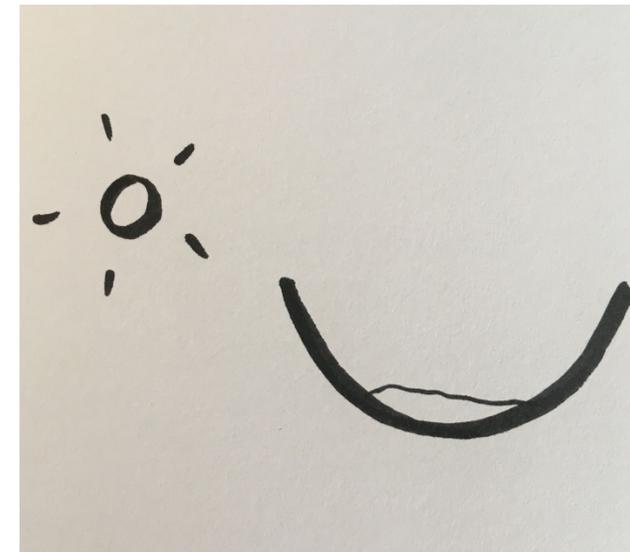
Step 2



Step 3



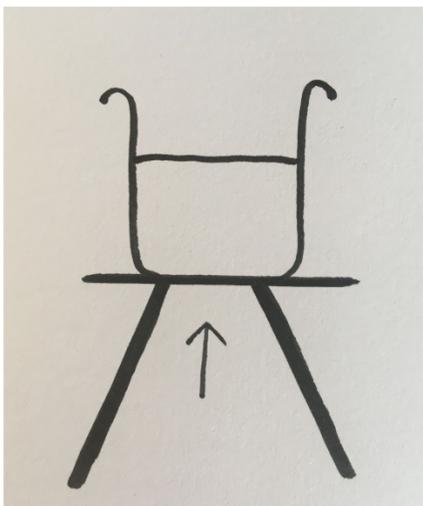
Step 4



Step 5

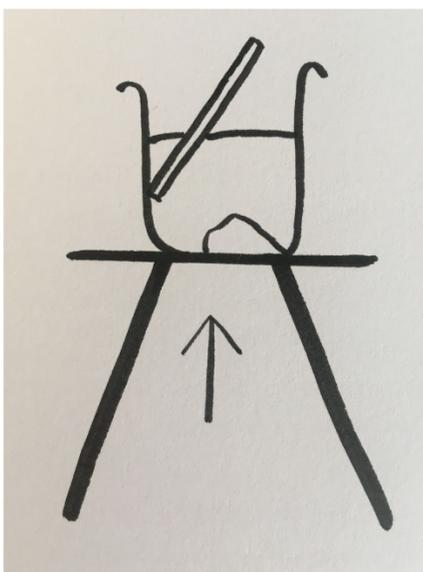


# Independent Practice



Step 1

Add the **sulfuric acid** to a **beaker** and heat gently using a **bunsen burner**. Be careful to make sure the acid doesn't boil and spit.



Step 2

Add the **copper oxide** and stir using a **glass rod**. Continue to gently heat (if needed) and keep adding copper oxide until there is excess.

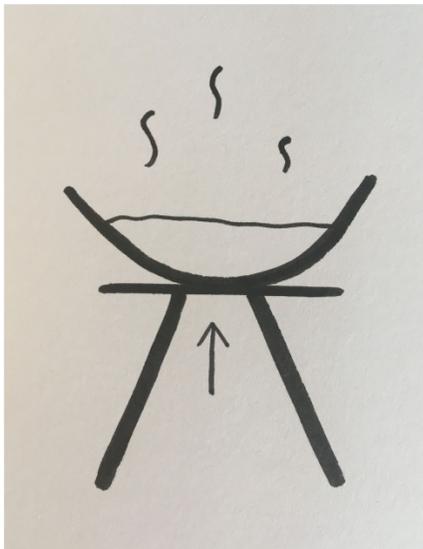


# Independent Practice



## Step 3

Put some **filter paper** into a **funnel** and place in a beaker. Pour the solution into the filter paper.

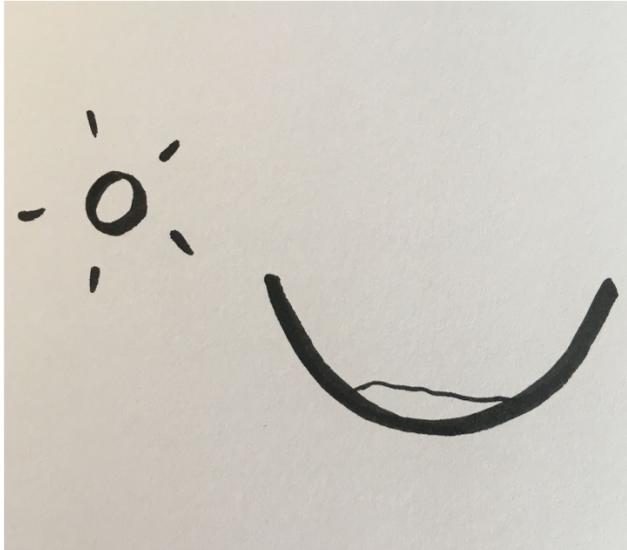


## Step 4

Pour the filtrate into an **evaporating basin** and place on a **tripod**. Heat gently to evaporate the water.



# Independent Practice



## Step 5

Once there is just a little bit of water left, leave on a windowsill to finish evaporating. Crystals of salt will remain.

