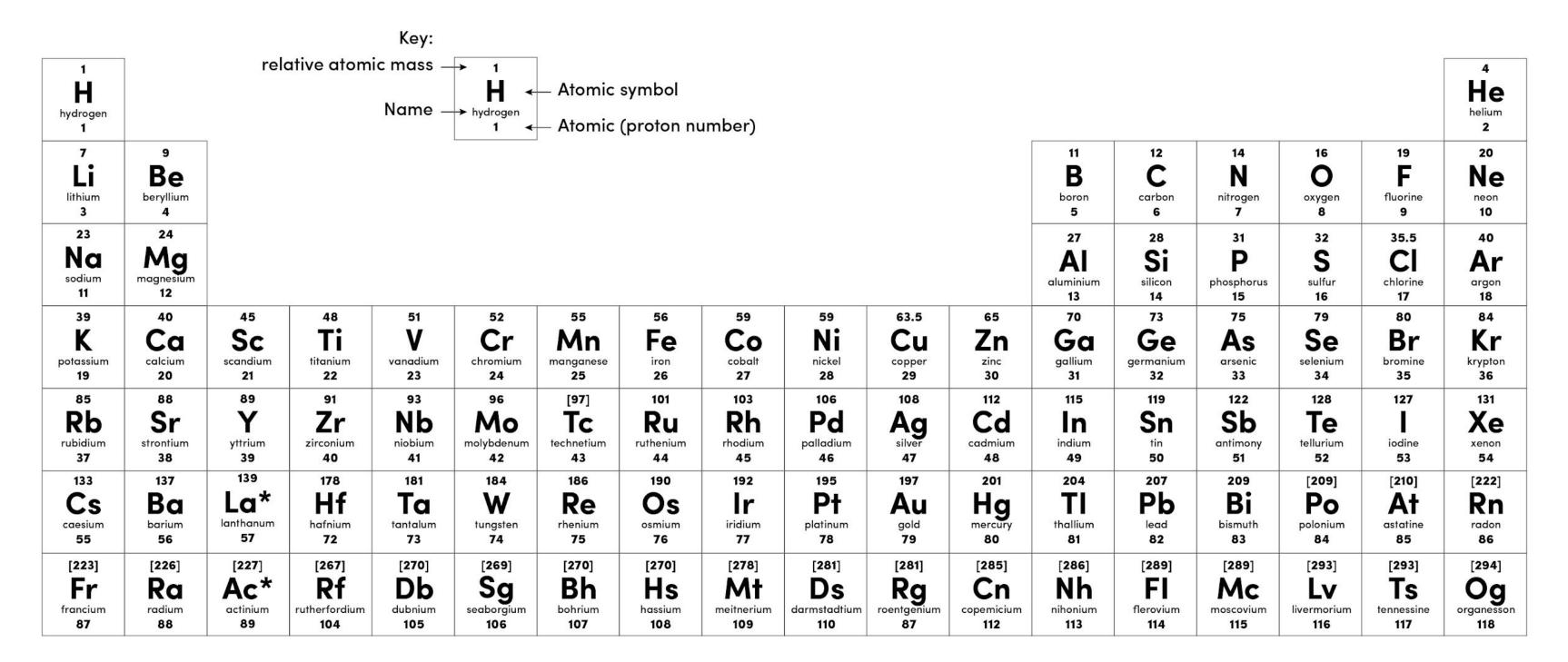
Combined Science - Chemistry - Key Stage 4 Atomic Structure & the Periodic Table

# Group 7 displacement reactions





### **Periodic Table of Elements**





#### Warm up

1. What name is given to group 7 elements?

2. How many electrons do group 7 elements have in their outer shell?

3. Do group 7 elements lose or gain electrons when they react?



## Independent practice

A student has obtained the results below.

	Fluorine	Chlorine	Bromine	Iodine
Potassium fluoride				
Potassium chloride				
Potassium bromide				
Potassium iodide				

- 1. Which halogen is the most reactive, and explain why using the student's results?
- 2. Which halide was bromine able to displace?



### Independent practice

For each reaction below **complete the equation**, and **state the colour change** that would be observed. If there is no reaction, state there is no reaction.

- 1. Sodium iodide + chlorine  $\rightarrow$
- 2. Sodium bromide + chlorine →
- 3. Sodium chloride + iodine →

#### Challenge:

Have a go at writing a balanced symbol equation for any reactions that occur

