

Reactivity

Lesson 14 - Displacement

Chemistry - Key Stage 3

Miss Fenner



Which is more reactive,
magnesium or tin?

Magnesium



Which is more reactive,
copper or iron?

Iron



Which is more reactive,
aluminium or calcium?

Calcium



Independent Practice

1. List the following metals from most to least reactive.
 - Copper
 - Magnesium
 - Iron
 - Tin
 - Gold
 - Sodium
2. How can you tell if a reaction is more vigorous?



Independent Practice

1. List the following metals from most to least reactive.
 - Sodium
 - Magnesium
 - Iron
 - Tin
 - Copper
 - Gold
2. When a reaction is more vigorous it will produced lots of **bubbles** quickly, change **colour** rapidly, produce a large change in **temperature** and take place **quickly**.



Will a reaction take place?

Iron + sodium oxide →

Option 1

Yes, absolutely, it will react!

Option 2

No, they can't react.



Will a reaction take place?

Tin + silver oxide →

Option 1

Yes, absolutely, it will react!

Option 2

No, they can't react.



Will a reaction take place?

Aluminium + zinc oxide →

Option 1

Yes, absolutely, it will react!

Option 2

No, they can't react.



Sodium + lead chloride → lead + sodium chloride



Calcium + tin oxide → tin + calcium oxide



Zinc + calcium fluoride → **NO REACTION!**



Independent Practice

Copy and complete these reactions. If no reaction will take place, write “no reaction”.

1. Lead + silver chloride →
2. Aluminium + sodium oxide →
3. Potassium + sodium fluoride →
4. Magnesium + zinc oxide →
5. Magnesium + calcium carbonate →



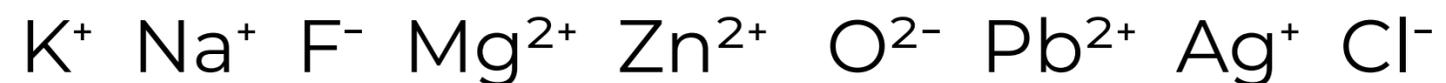
Independent Practice

1. Lead + silver chloride → silver + lead chloride
2. Aluminium + sodium oxide → no reaction
3. Potassium + sodium fluoride → sodium + potassium fluoride
4. Magnesium + zinc oxide → zinc + magnesium oxide
5. Magnesium + calcium carbonate → no reaction



Independent Practice

Write symbol equations for these reactions.



1. Potassium + sodium fluoride \rightarrow sodium + potassium fluoride
2. Magnesium + zinc oxide \rightarrow zinc + magnesium oxide
3. Lead + silver chloride \rightarrow silver + lead chloride
4. Sodium + magnesium oxide \rightarrow magnesium + sodium oxide



Independent Practice

Write symbol equations for these reactions.

